

## Standard State Thermodynamic Values At 298 15 K

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### Standard State Thermodynamic Values At

Standard-State Thermodynamic Values at 298.15 K. Standard-State Thermodynamic Values at 298.15 K: Enthalpy of Formation ( $\Delta H_f^\circ$ ), Free Energy of Formation ( $\Delta G_f^\circ$ ), and Absolute Entropy ( $S^\circ$ ) Substance  $\Delta H_f^\circ$  (kJ/mol. rxn)  $\Delta G_f^\circ$  (kJ/mol. rxn)  $S^\circ$ .

### Standard-State Thermodynamic Values at 298.15 K

Standard Thermodynamic Values at 25°C Please note that enthalpy and free energy values are given in kJ/mol while entropy values are given in J/(mol·K). Formula  $\Delta H_f^\circ$  (kJ/mol)  $\Delta G_f^\circ$  (kJ/mol)  $S^\circ$  (J/mol·K)  $\text{BOCl}_3$  (g) -1633.43 380.74 -1550.17  $\text{CN}_2$  (g) - cyanogen 308.95 242.25 297.19  $\text{NH}_2\text{CO}$  (s) - urea -333.51 104.60 -196.82  $\text{NH}_4$

### Standard Thermodynamic Values at 25°C - Chemistry-Reference

For a given material or substance, the standard state is the reference state for the material's thermodynamic state properties such as enthalpy, entropy, Gibbs free energy, and for many other material standards. The standard enthalpy change of formation for an element in its standard state is zero, and this convention allows a wide range of other thermodynamic quantities to be calculated and tabulated. The standard state of a substance does not have to exist in nature: for example, it is possible

### Standard state - Wikipedia

THERMODYNAMIC VALUES AT STANDARD STATE (298K) Data Retrieved From: Kots, Treichel, Weaver Chemistry & Chemical Reactivity (Sixth Edition) COPYRIGHT 2006! Species Name Enthalpy " $\Delta H^\circ$ " (kJ/mol) Entropy " $S^\circ$ " (J/(mol·K)) Gibbs energy " $\Delta G^\circ$ " (kJ/mol)  $\text{H}_2\text{O}$  (l) liquid water -285.83 69.95 -237.15  $\text{H}_2\text{O}$  (g) water vapor -241.83 188.84 -228.59 Al(s) Aluminum solid 0 28.3 0  $\text{AlCl}_3$ (s) Aluminum Chloride -705.63 109.29 -630.0  $\text{Al}_2\text{O}_3$ (s) AluminumOxide -1675.7 50.92 -1582.3  $\text{BaCl}_2$  (s) BariumChloride ...

### Thermodynamic Values at Standard State

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### Standard State Thermodynamic Values At 298 15 K ...

Standard Thermodynamic Values Formula State of Matter Enthalpy (kJ/mol) Entropy (J mol/K) Gibbs Free Energy (kJ/mol) (NH<sub>4</sub>)<sub>2</sub>O (l) -430.70096 267.52496 -267.10656 (NH<sub>4</sub>)<sub>2</sub>SiF<sub>6</sub> (s hexagonal) -2681.69296 280.24432 -2365.54992 (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (s) -1180.85032 220.0784 -901.90304 Ag (s) 0 42.55128 0 Ag (g) 284.55384 172.887064 245.68448

### Standard Thermodynamic Values - drjez.com

The standard state temperature is 25°C (298 K). It is possible to calculate standard state values for other temperatures. All liquids are pure. The concentration of all solutions is 1 M (1 molar). All gases are pure. All gases are at 1 atm pressure. The energy of formation of an element in its normal state is defined as zero.

### Standard State Conditions of Temperature and Pressure

\*Taken from "The NBS Tables of Chemical Thermodynamic Properties" (1982) and "CRC Handbook of Chemistry and Physics", 1st Student Edition (1988) ...

### Table of Thermodynamic Values - UW-Madison

Standard Thermodynamic Quantities for Chemical Substances at 25°C. Source of data: CRC Handbook of Chemistry and Physics, 84th Edition (2004). T1: Standard Thermodynamic Quantities - Chemistry LibreTexts

### T1: Standard Thermodynamic Quantities - Chemistry LibreTexts

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### Standard State Thermodynamic Values At 298 15 K

Standard state conditions are used for thermodynamic calculations. Several conditions are specified for the standard state: The standard state temperature is 25 degrees C (298 K). Note that temperature is not specified for standard state conditions, but most tables are compiled for this temperature.

### Standard Conditions Versus Standard State

Standard Thermodynamic Values at 25.0°C (298.15 K)

### Standard Thermodynamic Values - The Art of Chemistry

In chemistry, the standard state of a material is its state at 1 bar (100 kilopascals exactly). This pressure was changed from 1 atm (101.325 kilopascals) by IUPAC in 1990.

### Standard\_state - chemeurope.com

Table of Contents. This page contains several tables detailing the standard thermodynamic properties for several different substances. The table has been separated by substance, as listed below:

### Standard Thermodynamic Properties for Selected Substances ...

Thermodynamic databases contain information about thermodynamic properties for substances, the most important being enthalpy, entropy, and Gibbs free energy. Numerical values of these thermodynamic properties are collected as tables or are calculated from thermodynamic datafiles. Data is expressed as temperature-dependent values for one mole of substance at the standard pressure of 101.325 kPa, or 100 kPa. Unfortunately, both of these definitions for the standard condition for pressure are in use.

### Thermodynamic databases for pure substances - Wikipedia

The properties tabulated are:  $\Delta_f H^\circ$  Standard molar enthalpy (heat) of formation at 298.15 K in kJ/mol  $\Delta_f G^\circ$  Standard molar Gibbs energy of formation at 298.15 K in kJ/mol  $S^\circ$  Standard molar entropy at 298.15 K in J/mol K  $C_p$  Molar heat capacity at constant pressure at 298.15 K in J/mol K The standard state pressure is 100 kPa (1 bar).

### STANDARD THERMODYNAMIC PROPERTIES OF CHEMICAL SUBSTANCES

Standard Thermodynamic Values at 25°C Please note that enthalpy and free energy values are given in kJ/mol while entropy values are given in J/(mol·K).

### Standard Thermodynamic Values at 25°C - Chemistry-Reference

Once the values for all the reactants and products are known, the standard Gibbs free energy change for the reaction can be found. Most tables of thermodynamic values list  $\Delta_f G^\circ$ 's for common substances. They can, of course, always be found from values of  $\Delta_f H^\circ$ ,  $\Delta_f H^\circ$ , and  $\Delta_f S^\circ$ .

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