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Problem 2 In an
experiment to identify

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an unknown gas, it is found that 1.82 L of the gas has a mass of 5.430 g. What is the density of the gas in g/L? Solution ANALYZE What is given in the problem? the measured mass and volume of the gas

Skills Worksheet Problem Solving

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Mole Concept Name
Class Date Problem

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Solving continued

CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS

Sample Problem 4 uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

Skills Worksheet

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Problem Solving

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Answer
Titrations Titrations

Chemists have many methods for determining the quantity of a substance present in a solution or other mixture. One common method is titration, in which a solution of known concentration reacts with a sample containing the substance of unknown

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quantity.

Skills Worksheet
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Conversions

Conversions One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how

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much, how big, how
hot, how many, how
hard, and how long did
it take.

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Mole Concept Name

Class Date • Problem

Solving continued

PROBLEMS INVOLVING
ATOMS AND ELEMENTS

Sample Problem 1 A

chemist has a jar

containing 388.2 g of

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iron filings. How many moles of iron does the jar contain? Solution
ANALYZE What is given in the problem? mass of iron in grams

continued - PC\|MAC

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pH Name Class Date

Problem Solving

continued Sample

Problem 1 A HCl

solution has a

concentration of

0.0050 M. Calculate

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[OH⁻] and [H₃O⁺] for this solution. HCl is a strong acid, so assume it is 100% ion-ized.

Solution ANALYZE

What is given in the problem? the molarity of the HCl solution, and the fact

Skills Worksheet Problem Solving

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Stoichiometry Name

Class Date Problem

Solving continued

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Sample Problem 1

Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$. If 4.0 mol of H_2 react, how many moles of NH_3 will be produced?

Skills Worksheet **Problem Solving**

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Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they relate moles of a substance to mass by using the molar mass.

Skills Worksheet Problem Solving - Mole Cafe

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Conversions

Conversions One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big, how hot, how many, how hard, and how long did it take.

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Concentration of
Solutions Sample
Problem 3 Determine
the molal
concentration of a
solution containing
81.3 g of ethylene
glycol, HOCH₂CH₂
OH, dissolved in 166 g
of water. Solution

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ANALYZE What is given in the problem? the mass of ethylene glycol dissolved, and What are you asked to find?
Items Data

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Mole Concept Name

Class Date Problem

Solving continued Is
the answer

reasonable? Yes; 2 g of
boron is about $\frac{1}{5}$ of

the molar mass of
boron. Therefore, 2.00

g boron will contain

about $\frac{1}{5}$ of an

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Avogadro's constant of
atoms. Practice 1.

Answer

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Stoichiometry Name

Class Date Problem

Solving continued

General Plan for

Solving Stoichiometry

Problems Convert

using the mole ratio A,

given in the balanced

chemical equation

Mass of substance A

Amount in mol of

substance A Amount in

mol of substance B

Convert using the

molar mass of A 1

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